

Honors Chemistry 2

The following polyatomic ions should be memorized:

+1 ions		-1 ions		-2 ions	
$\text{NH}_4^+$	ammonium	$\text{CH}_3\text{COO}^-$	acetate	$\text{CO}_3^{2-}$	carbonate
		$\text{C}_2\text{H}_3\text{O}_2^-$	acetate	$\text{CrO}_4^{2-}$	chromate
+2 ions		$\text{BrO}_3^-$	bromate	$\text{Cr}_2\text{O}_7^{2-}$	dichromate
$\text{Hg}_2^{2+}$	Mercury (I)	$\text{ClO}_3^-$	chlorate	$\text{C}_2\text{O}_4^{2-}$	oxalate
		$\text{IO}_3^-$	iodate	$\text{O}_2^{2-}$	peroxide
-3 ions		$\text{NO}_3^-$	nitrate	$\text{SO}_4^{2-}$	sulfate
$\text{PO}_4^{3-}$	phosphate	$\text{MnO}_4^-$	permanganate	$\text{SiO}_3^{2-}$	silicate
		$\text{OH}^-$	hydroxide	$\text{S}_2\text{O}_3^{2-}$	thiosulfate
		$\text{CN}^-$	cyanide		

All derivatives of the above are also to be known; examples follow

$\text{BrO}_4^-$	Perbromate	$\text{SO}_5^{2-}$	Persulfate	$\text{HPO}_4^{2-}$	Hydrogen Phosphate
$\text{BrO}_2^-$	Bromite	$\text{SO}_3^{2-}$	Sulfite	$\text{H}_2\text{PO}_4^-$	Dihydrogen phosphate
$\text{BrO}^-$	Hypobromite	$\text{SO}_2^{2-}$	Hyposulfite	$\text{PO}_3^{3-}$	Phosphite
$\text{ClO}_4^-$	Perchlorate	$\text{HSO}_4^-$	Bisulfate		
$\text{ClO}_2^-$	Chlorite	$\text{HSO}_4^-$	Hydrogen sulfate		
$\text{ClO}^-$	hypochlorite	$\text{HSO}_3^-$	Bisulfite		
$\text{IO}_4^-$	Periodate	$\text{HSO}_2^-$	Hydrogen sulfite		
$\text{IO}_2^-$	Iodite	$\text{HCO}_3^-$	Bicarbonate		
$\text{IO}^-$	Hypoiodite	$\text{HCO}_3^-$	Hydrogen carbonate		
$\text{NO}_4^-$	Pernitrate	$\text{CrO}_3^{2-}$	chromite		
$\text{NO}_2^-$	Nitrite				
$\text{NO}^-$	Hyponitrite				