

# Things to Memorize

## I. Polyatomic Ions

1 <sup>+</sup> ions		1 <sup>-</sup> ions		2 <sup>-</sup> ions		3 <sup>-</sup> ions	
NH <sub>4</sub> <sup>+</sup>	Ammonium	BrO <sub>3</sub> <sup>-</sup>	Bromate	SO <sub>4</sub> <sup>2-</sup>	Sulfate	PO <sub>4</sub> <sup>3-</sup>	Phosphate
		IO <sub>3</sub> <sup>-</sup>	Iodate	SO <sub>3</sub> <sup>2-</sup>	Sulfite		
		ClO <sub>3</sub> <sup>-</sup>	Chlorate	CrO <sub>4</sub> <sup>2-</sup>	Chromate		
		ClO <sub>2</sub> <sup>-</sup>	Chlorite	CO <sub>3</sub> <sup>2-</sup>	Carbonate		
		NO <sub>3</sub> <sup>-</sup>	Nitrate	SiO <sub>3</sub> <sup>2-</sup>	Silicate		
		NO <sub>2</sub> <sup>-</sup>	Nitrite				
		OH <sup>-</sup>	Hydroxide				
		CN <sup>-</sup>	Cyanide				

## II. Common Acids

HCl	Hydrochloric Acid
HF	Hydrofluoric Acid
HI	Hydroiodic Acid
HBr	Hydrobromic Acid
H <sub>2</sub> SO <sub>4</sub>	Sulfuric Acid
H <sub>2</sub> SO <sub>3</sub>	Sulfurous Acid
HNO <sub>3</sub>	Nitric Acid
HNO <sub>2</sub>	Nitrous Acid

## III. Common Ions (Latin suffixes)

Fe <sup>2+</sup>	Ferrous
Fe <sup>3+</sup>	Ferric
Cu <sup>+</sup>	Cuprous
Cu <sup>2+</sup>	Cupric
Sn <sup>2+</sup>	Stannous
Sn <sup>4+</sup>	Stannic
Pb <sup>2+</sup>	Plumbous
Pb <sup>4+</sup>	Plumbic