VALENCE ELECTRONS

Name _____

The valence electrons are the electrons in the outermost principal energy level. They are always "s" or "s and p" electrons. Since the total number of electrons possible in s and p sublevels is eight, there can be no more than eight valence electrons.

Determine the number of valence electrons in the atoms below.

Example: carbon

Electron configuration is 1s² | 2s² 2p² |

Carbon has 4 valence electrons.

1. fluorine _____

11. lithium _____

2. phosphorus _____

12. zinc _____

3. calcium _____

13. carbon _____

4. nitrogen _____

14. iodine _____

5. iron _____

15. oxygen _____

6. argon _____

16. barium _____

7. potassium _____

17. aluminum _____

8. helium

18. hydrogen _____

9. magnesium _____

19. xenon _____

10. sulfur _____

20. copper _____