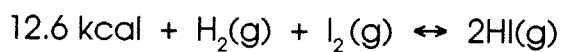
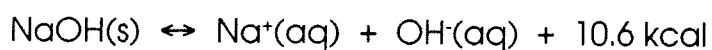


LE CHATELIER'S PRINCIPLE CONTINUED

Name _____



| Stress | Equilibrium Shift | [H ₂] | [I ₂] | [HI] | K |
|--------------------------|-------------------|-------------------|-------------------|-----------|------------------|
| 1. Add H ₂ | right | _____ | decreases | increases | remains the same |
| 2. Add I ₂ | | | _____ | | |
| 3. Add HI | | | | _____ | |
| 4. Remove H ₂ | | _____ | | | |
| 5. Remove I ₂ | | | _____ | | |
| 6. Remove HI | | | | _____ | |
| 7. Increase Temperature | | | | | |
| 8. Decrease Temperature | | | | | |
| 9. Increase Pressure | | | | | |
| 10. Decrease Pressure | | | | | |



(Remember that pure solids and liquids do not affect equilibrium values.)

| Stress | Equilibrium Shift | Amount NaOH(s) | [Na ⁺] | [OH ⁻] | K |
|---|-------------------|----------------|--------------------|--------------------|---|
| 1. Add NaOH(s) | | _____ | | | |
| 2. Add NaCl (Adds Na ⁺) | | | _____ | | |
| 3. Add KOH (Adds OH ⁻) | | | | _____ | |
| 4. Add H ⁺ (Removes OH ⁻) | | | | _____ | |
| 5. Increase Temperature | | | | | |
| 6. Decrease Temperature | | | | | |
| 7. Increase Pressure | | | | | |
| 8. Decrease Pressure | | | | | |